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1 a) The relative atomic mass of an element is the weighted average mass of the isotopes of the element. (1) It is measured on a scale on which a carbon-12 atom has a mass of exactly 12 units. (1). b) The total

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mass of 100 atoms =

$$(75 \times 35) + (25 \times 37)$$

$$= 3550 \text{ (1) The}$$

average mass of 1

$$\text{atom} = 3550/100 =$$

$$35.5 \text{ (1)}$$

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sublimation d)

subliming /

sublimation 2. a)

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have regularly packed  
particles touching.

Liquids should have  
most of the

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